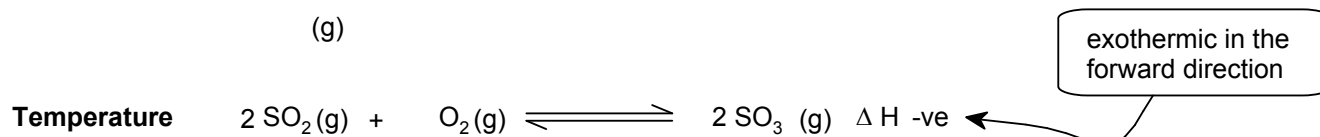
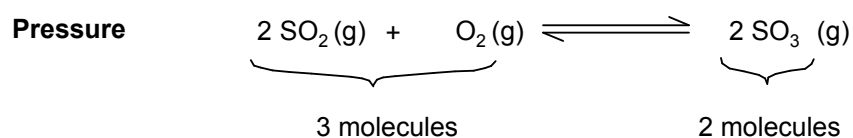


Equilibrium and Le Chatelier's Principle

| | | |
|-----------------|-------------------|---------------------------------------|
| Temperature ↑ | equilibrium moves | in the endothermic direction |
| Pressure ↑ | equilibrium moves | towards the side with fewer molecules |
| Concentration ↑ | equilibrium moves | to absorb the species |



If the temperature is increased, the equilibrium will move in the endothermic direction.
i.e. towards the left hand side.



If the pressure is increased, the equilibrium will move towards the side with fewer molecules.
i.e. towards the right hand side.



If the concentration of SO_2 is increased, the equilibrium will move to absorb SO_2 .
i.e. towards the right hand side.