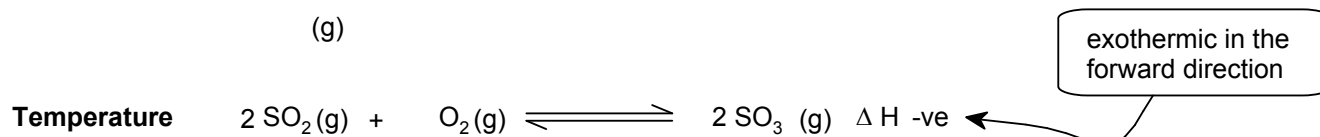
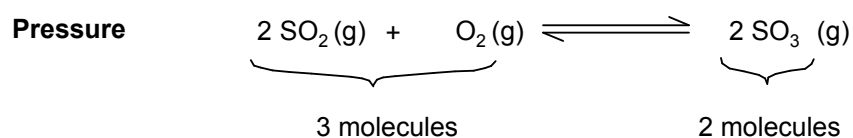


Equilibrium and Le Chatelier's Principle

Temperature ↑	equilibrium moves	in the endothermic direction
Pressure ↑	equilibrium moves	towards the side with fewer molecules
Concentration ↑	equilibrium moves	to absorb the species



If the temperature is increased, the equilibrium will move in the endothermic direction.
i.e. towards the left hand side.



If the pressure is increased, the equilibrium will move towards the side with fewer molecules.
i.e. towards the right hand side.



If the concentration of SO_2 is increased, the equilibrium will move to absorb SO_2 .
i.e. towards the right hand side.